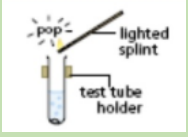
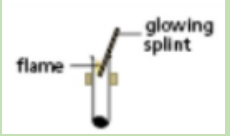
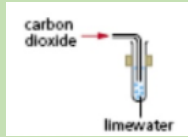




Chemistry: Reactions

Key word	Definition
Chemical reaction	Atoms (reactants) rearrange themselves and join back together in a different way to make the products.
Fuels	Material such as coal, gas, or oil that are burned to produce heat or power.
Exothermic reaction	A reaction where the temperature increases
Endothermic reactions	A reaction where the temperature decreases

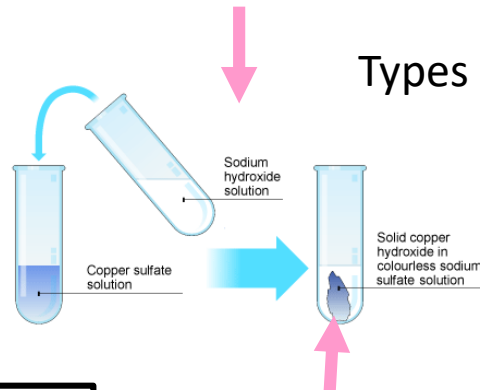
Testing for gases	
Hydrogen	
Oxygen	
Carbon Dioxide	

Type of Change	Description	Example
Chemical	<ul style="list-style-type: none"> Irreversible Re-arranges atoms (chemical reaction) Change properties 	Combustion Exothermic reaction 
Physical	<ul style="list-style-type: none"> Reversible Atoms not re-arranged (no chemical reaction) Same properties 	Change of state 

The **Fire Triangle** or **Combustion Triangle** shows the three elements a **fire** needs to ignite: heat, fuel, and oxygen.

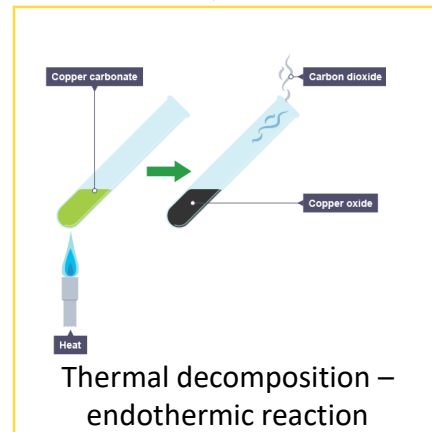


A precipitation reaction is where two liquids are mixed together to form a **solid substance**



The solid is the precipitate

A decomposition reaction is where a compound breaks down into simpler compounds or elements



Chemical equations:
Reactants → **Products**
 Copper + Oxygen → Copper Oxide

Reactivity series

- potassium **most reactive**
- sodium
- calcium
- magnesium
- aluminium
- carbon
- zinc
- iron
- tin
- lead
- hydrogen
- copper
- silver
- gold
- platinum **least reactive**
- 